

DOI: 10.1515/amm-2017-0292

A. DĘBSKI*, B. ONDERKA**, W. GAŚSIOR*[#], T. GANCARZ***PHASE EQUILIBRIA IN THE Bi-In-Sn-Zn SYSTEM. THERMAL ANALYSIS VS. CALCULATIONS**

With the use of the differential thermal analysis (DTA), studies of the phase transitions were conducted for 90 of alloys from the quaternary Bi-In-Sn-Zn system and for the constant ratio of Bi:In and Bi:Sn. The studies were conducted for the alloys prepared from the purity metals (Bi, In, Sn, Zn = 99.999 mas. %) by way of melting in a graphite crucible in a glove-box filled with Ar, in which the impurities level was less than 0.1 ppm. After melting and thorough mixing, the liquid alloys were poured out into a graphite test mold. The phase transition temperature data obtained from the DTA investigations were next confronted with those determined from the calculations based on the binary and ternary optimized thermodynamic parameters available in the literature. It was found that the experimental and the calculated phase transition temperatures were in good agreement.

Keywords: Thermodynamics, Differential Thermal Analysis, Quaternary Alloys

1. Introduction

The quaternary Bi-In-Sn-Zn system is important for the lead-free solder technology because its Sn-rich alloys are potential substitutes of traditional lead solders based on the near eutectic Pb-Sn alloys which have been used for many years in microelectronics. The first experimental data on the wettability of Cu-substrates by Bi-In-Sn-Zn solders of a high Sn concentration were very promising [1] because of their low melting temperature being close to that for the Pb-Sn eutectic 456 K (183°C), and their better mechanical properties, in particular the creep resistance and the tensile strength.

The improvement of the different properties of these lead-free solders needs a systematic updating of the thermo-physical properties as well as of the thermodynamic data base for the critical evaluation of all the experimental data available in the literature by means of the CALPHAD (**C**alculation of **A**lloys **P**hase **D**iagrams) method.

This procedure is based on sequential modeling of the phase equilibria and thermodynamic properties of alloy systems, starting from binary systems and extending to higher-order systems, to obtain accurate predictions of the phase equilibria in multicomponent systems. The crucial condition for the success of such a method is a well-updated CALPHAD database for the calculation of the thermodynamic properties of the relevant systems. The accuracy of such a database relies on the accuracy of the interaction energies assessed for the binary and the ternary systems. Ideally, these energies can be extracted from the available experimental data by applying different models.

Only limited information on the quaternary Bi-In-Sn-Zn system is available. Some Sn-rich alloys have been analyzed by differential scanning calorimetry [2] and their microstructure was examined with the use of SEM (Scanning Electron Microscopy) and XRD (X-Ray Diffraction). This is, basically, the only available data, as no extensive investigations of this type of quaternary alloys have been conducted, so far.

The Bi-In-Sn-Zn system was widely assessed by Moelans et al. [3], who used all the available thermodynamic data and assessed the thermodynamic parameters of the constituent binary systems. Using the CALPHAD method, they optimized the thermodynamic descriptions for the ternary Bi-In-Sn, Bi-In-Zn, Bi-Sn-Zn and In-Sn-Zn systems, and combined them to obtain a description of the quaternary Bi-In-Sn-Zn system. Because Moelans et al. [3] provide no information on the thermodynamic properties and topology of the Bi-In-Zn system, an approximate description was elaborated to reproduce the liquid miscibility gap with a plausible shape. In the present work, the new assessment of the Bi-In-Zn system [4] was adopted, which was performed with the use of the new experimental data: the DTA analysis of 15 samples by the DTA technique in the temperature range of 300-900 K (27-627°C) [4], the enthalpy of mixing of liquid alloys [5] determined at 773 K (500°C) and 873 K (600°C) by the drop calorimetric technique in a Calvet-type micro-calorimeter and the zinc activity [6] of the liquid Bi-In-Zn solutions obtained from the e.m.f. (LiCl+KCl+Zn electrolyte) method.

In this work, DTA studies of the transition temperatures were conducted for 90 of Bi-In-Sn-Zn alloys for selected ratios of the mole concentrations. These calculations were conducted

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based on the interaction parameters for the binary and ternary systems available in the literature [4,7]. Therefore, the thermodynamic database of quaternary Bi-In-Sn-Zn system was combined from assessments of four constituent ternary systems. It is a common practice for multicomponent systems because of lack of thermodynamic data. The control experiments (for example DTA) allow to check the accuracy of such description. For this purpose the DTA experimental data and those obtained from the phase equilibria calculations were compared. Otherwise, the present work can be treated as such check of coherency of optimized ternary parameters.

2. DTA study of Bi-In-Sn-Zn system

2.1. Alloys preparation

The Bi-In-Sn-Zn alloys were prepared from metals: In 99.999 mass. % (Indium Corporation, USA), Bi, Sn and Zn all 99.999 mass. % (ITME, Warsaw, Poland). The weighed amounts of pure metals were melted in graphite crucibles in a glove box under purified Ar (impurities < 0.1 ppm). After heating up to 773 K (500°C) and mixing of the components, the resulting alloys were cast into a graphite sample mold. All together 90 alloys with Zn concentration changing from 0.05 to 0.9 mole fraction and the different ratios of the remaining components: Bi:In (0.25, 0.5 and 0.75) and Bi:Sn (0.067, 0.111, 0.143, 0.2, 0.333, 0.428, 0.5, 0.6, 1.286) were produced. The concentrations of all the alloys applied in our studies are presented in the next part, together with the results of the phase transition temperatures from the DTA studies.

2.2. Apparatus

The Bi-In-Sn-Zn system was studied by the DTA method with the use of a modified Q-1500 D (Paulik-Paulik-Erdey) (MOM, Hungary) apparatus working up to 1300 K (1027°C). Based on the many calibration measurements of the pure metals, the accuracy of the DTA results is less than ± 2 K for Q-1500 D. The alloy samples of about 1 gram were placed into alumina crucibles after the division of the cast material into 2 parts. Using Q-1500 D the DTA heating and cooling cycle rate was 5 deg/min in the temperature range from 300 up to max. 900 K (27-627 °C) limit taking into account the temperature of the miscibility gap boundary. And so, most of the alloys were investigated at temperatures lower than 700 K (427 °C). To avoid a concentration change of the alloys due to zinc evaporation, the DTA investigations were performed only in one heating and cooling cycle. In comparison to the partial pressure of Zn, the partial pressures of Bi and In at temperatures lower than 900 K (627 °C) are several orders of magnitude lower. Therefore, their contents in the samples should not change considerably during the experiments.

Additionally, to test the reliability of the presented thermal analysis results, DTA measurements were performed with the use

of and DSC/DTA Pegasus 404 F1 (Netzsch, Germany) scanning calorimeter with a DTA probe, in a similar temperature range. According to the recommendations of Boettinger et al. [10] the univariant and invariant transitions were determined from onset temperatures and the liquidus from the last peak of DTA signal. The samples groups of numbers (23, 25, 27), (51, 53, 57) and (87, 89) were analyzed in 40-770 K (-233-497°C) temperature range with the heating and cooling rates of 2 deg/min in two consecutive cycles. Also the samples no. 37, 38 and 68 were measured with the heating and cooling rates of 2 deg/min and 5 deg/min. The apparatus was calibrated for the melting point of pure metals (min. 99.999 mas.%): In, Sn, Zn, Pb, Sb and Ag. The samples were sealed in evacuated silica tubes and placed at a DTA sensor in a platinum crucible. An empty evacuated silica tube in a platinum crucible was used as a reference state. No traces of Zn evaporation was observed at inner surface of quartz tubes.

In the following publication the abbreviations Q-1500 D and Netzsch 404 DTA for the DTA devices were used.

2.3. Results

The DTA studies were performed by Q-1500 D in two series of heating and cooling with the heating rate equaling 5 deg/min, and the obtained results are shown in Table 1a-i and in Figs 1a-i. The concentrations of the components in the Tables are in atomic percentage.

TABLE 1a

The DTA results of Bi-In-Sn-Zn alloys.
 Cross sections: (Bi:In = 0.25 and Bi:Sn = 0.6) + % at. Zn

Sample no.	Bi % at.	In % at.	Sn % at.	Zn % at.	Registered DTA thermal arrests, °C
1	14.25	57.00	23.75	5.00	61; 67
2	13.50	54.00	22.50	10.00	61; 69
3	12.00	48.00	20.00	20.00	61; 67; 290*
4	10.50	42.00	17.50	30.00	61; 68; 324
5	9.00	36.00	15.00	40.00	63; 354
6	7.50	30.00	12.50	50.00	64; 366
7	6.00	24.00	10.00	60.00	51; 89; 357
8	4.50	18.00	7.50	70.00	60; 362
9	3.00	12.00	5.00	80.00	60; 365
10	1.50	6.00	2.50	90.00	60; 371

* Weak thermal signals.

For some alloys, the signals at the DTA diagram were observed only in one cycle (cooling or heating) and they were designated in Table 1a-i by a question mark. It should also be noted that, for the same samples, even for the cooling velocity equaling 2 deg/min, an undercooling of about 10°C was observed. Very weak thermal signals were also registered, which were not observed in the other samples or the thermal cycles. These results were additionally marked in Table 1a-i by an asterisk.

TABLE 1b

The DTA results of Bi-In-Sn-Zn alloys.
 Cross sections: (Bi:In = 0.25 and Bi:Sn = 0.2) + % at. Zn

Sample no.	Bi % at.	In % at.	Sn % at.	Zn % at.	Registered DTA thermal arrests, °C
11	9.50	38.00	47.50	5.00	60; 115
12	9.00	36.00	45.00	10.00	60; 115; 197*
13	8.00	32.00	40.00	20.00	60; 109; 120; 261
14	7.00	28.00	35.00	30.00	60; 115; 299
15	6.00	24.00	30.00	40.00	61; 72; 121; 331
16	5.00	20.00	25.00	50.00	61; 73; 123; 351
17	4.00	16.00	20.00	60.00	51; 115; 351
18	3.00	12.00	15.00	70.00	51; 118; 357
19	2.00	8.00	10.00	80.00	53; 112; 355
20	1.00	4.00	5.00	90.00	373

* Weak thermal signals.

TABLE 1e

The DTA results of Bi-In-Sn-Zn alloys.
 Cross sections: (Bi:In = 0.5 and Bi:Sn = 0.33) + % at. Zn

Sample no.	Bi % at.	In % at.	Sn % at.	Zn % at.	Registered DTA thermal arrests, °C
41	15.83	31.67	47.50	5.00	60; 124
42	15.00	30.00	45.00	10.00	60; 124
43	13.33	26.67	40.00	20.00	60; 124
44	11.67	23.33	35.00	30.00	61; 125; 311
45	10.00	20.00	30.00	40.00	63; 124; 340
46	8.33	16.67	25.00	50.00	70; 124; 359
47	6.67	13.33	20.00	60.00	57; 118; 356
48	5.00	10.00	15.00	70.00	65; 117; 363
49	3.33	6.67	10.00	80.00	60; 124; 368
50	1.67	3.33	5.00	90.00	375

TABLE 1c

The DTA results of Bi-In-Sn-Zn alloys.
 Cross sections: (Bi:In = 0.25 and Bi:Sn = 0.067) + % at. Zn

Sample no.	Bi % at.	In % at.	Sn % at.	Zn % at.	Registered DTA thermal arrests, °C
21	4.75	19.00	71.25	5.00	131; 166
22	4.50	18.00	67.50	10.00	131; 174
23	4.00	16.00	60.00	20.00	125; 160; 240
24	3.50	14.00	52.50	30.00	126; 160; 279
25	3.00	12.00	45.00	40.00	146; 311
26	2.50	10.00	37.50	50.00	148; 338
27	2.00	8.00	30.00	60.00	155; 339
28	1.50	6.00	22.50	70.00	155; 347
29	1.00	4.00	15.00	80.00	156; 365
30	0.50	2.00	7.50	90.00	156; 373

TABLE 1f

The DTA results of Bi-In-Sn-Zn alloys.
 Cross sections: (Bi:In = 0.5 and Bi:Sn = 0.11) + % at. Zn

Sample no.	Bi % at.	In % at.	Sn % at.	Zn % at.	Registered DTA thermal arrests, °C
51	7.92	15.83	71.25	5.00	63; 153; 166
52	7.50	15.00	67.50	10.00	63; 161; 179
53	6.67	13.33	60.00	20.00	161; 244
54	5.83	11.67	52.50	30.00	161; 288
55	5.00	10.00	45.00	40.00	321
56	4.17	8.33	37.50	50.00	327
57	3.33	6.67	30.00	60.00	345
58	2.50	5.00	22.50	70.00	354
59	1.67	3.33	15.00	80.00	365
60	0.83	1.67	7.50	90.00	374

TABLE 1d

The DTA results of Bi-In-Sn-Zn alloys.
 Cross sections: (Bi:In = 0.5 and Bi:Sn = 1) + % at. Zn

Sample no.	Bi % at.	In % at.	Sn % at.	Zn % at.	Registered DTA thermal arrests, °C
31	23.75	47.50	23.75	5.00	60; 73
32	22.50	45.00	22.50	10.00	62; 75; 232*
33	20.00	40.00	20.00	20.00	66; 77
34	17.50	35.00	17.50	30.00	67; 339*
35	15.00	30.00	15.00	40.00	60; 368
36	12.50	25.00	12.50	50.00	66; 377
37	10.00	20.00	10.00	60.00	55; 367; 394
38	7.50	15.00	7.50	70.00	59; 367; 419
39	5.00	10.00	5.00	80.00	60; 374
40	2.50	5.00	2.50	90.00	59; 380

* Weak thermal signals.

TABLE 1g

The DTA results of Bi-In-Sn-Zn alloys.
 Cross sections: (Bi:In = 0.75 and Bi:Sn = 1.286) + % at. Zn

Sample no.	Bi % at.	In % at.	Sn % at.	Zn % at.	Registered DTA thermal arrests, °C
61	30.54	40.71	23.75	5.00	60?; 73; 186
62	28.93	38.57	22.50	10.00	60; 73; 243
63	25.71	34.29	20.00	20.00	61?; 73; 312
64	22.50	30.00	17.50	30.00	61?; 74; 355
65	19.29	25.71	15.00	40.00	65; 263; 340
66	16.07	21.43	12.50	50.00	65; 373
67	12.86	17.14	10.00	60.00	76; 376
68	9.64	12.86	7.50	70.00	75; 302; 375
69	6.43	8.57	5.00	80.00	60; 80; 371
70	3.21	4.29	2.50	90.00	60; 78; 381

? Thermal arrest detected for cooling.

The results obtained from additional DTA measurements (Netzsch 404 DTA) for selected samples are given in Table 2a-b and were superimposed in Figs 1c-d, Figs 1f-g, and Fig. 1i, as the mean values of the transition temperatures determined based on two heating and cooling cycles. The difference between the results of the onset temperatures of two heating cycles was lower

than $\pm 2^\circ\text{C}$ (Figs 2a-b). In the case of sample 68 (2 deg/min heating cycle – Table 2b) a significant difference between the beginning of the signal change (b) and the onset temperature of about $\sim 4.8^\circ\text{C}$ can be observed (Fig. 3.). In the case of doubled cycles acquired with different cooling rates the difference between the peak endset temperatures was about $\pm 0.5^\circ\text{C}$.

TABLE 1h

The DTA results of Bi-In-Sn-Zn alloys.
Cross sections: (Bi:In = 0.75 and Bi:Sn = 0.43) + % at. Zn

Sample no.	Bi % at.	In % at.	Sn % at.	Zn % at.	Registered DTA thermal arrests, °C
71	20.36	27.14	47.50	5.00	61; 80; 125; 376*
72	19.29	25.71	45.00	10.00	78; 121; 196
73	17.14	22.86	40.00	20.00	125; 273
74	15.00	20.00	35.00	30.00	63; 125; 316
75	12.86	17.14	30.00	40.00	63; 125; 346
76	10.71	14.29	25.00	50.00	62; 125; 364
77	8.57	11.43	20.00	60.00	75; 125; 266; 332; 360*
78	6.43	8.57	15.00	70.00	75; 125; 369
79	4.29	5.71	10.00	80.00	78; 125; 372
80	2.14	2.86	5.00	90.00	75; 125; 376

* Weak thermal signals.

TABLE 1i

The DTA results of Bi-In-Sn-Zn alloys.
Cross sections: (Bi:In = 0.75 and Bi:Sn = 0.143) + % at. Zn

Sample no.	Bi % at.	In % at.	Sn % at.	Zn % at.	Registered DTA thermal arrests, °C
81	10.18	13.57	71.25	5.00	61; 164
82	9.64	12.86	67.50	10.00	66; 160; 182
83	8.57	11.43	60.00	20.00	61; 161; 246
84	7.50	10.00	52.50	30.00	64; 160; 289
85	6.43	8.57	45.00	40.00	62; 161; 319
86	5.36	7.14	37.50	50.00	63; 161; 345
87	4.29	5.71	30.00	60.00	78; 344
88	3.21	4.29	22.50	70.00	78; 357
89	2.14	2.86	15.00	80.00	147?; 363
90	1.07	1.43	7.50	90.00	145?; 378

? Thermal arrest detected at cooling.

TABLE 2a

The results of DTA experiments for quaternary Bi-In-Sn-Zn alloys of chosen compositions acquired by DTA/DSC Pegasus 404 F1 (Netzsch, Germany)

Sample no.	Bi at. %	In at. %	Sn at. %	Zn at. %	Thermal arrests*, °C
Heating (H) and cooling (C) – rate 2 deg/min					
(Fig. 1c)					
23	4.00	16.00	60.00	20.00	58.5; 168.8; 270.4
					57.4; 165.8; 241.7
25	3.00	12.00	45.00	40.00	58.2; 151.8; 168.2; 325.5
					56.1; 163.0; 353.5
27	2.00	8.00	30.00	60.00	58.3; 148.2; 167.5; 360.2
					56.1; 162.9; 353.8
(Fig. 1f)					
51	7.92	15.83	71.25	5.00	77.8; 162.4; 170.1
					76.2; 155.0; 168.1
53	6.67	13.33	60.00	20.00	77.7; 167.7; 254.5
					163.1; 247.9
57	3.33	6.67	30.00	60.00	78.9; 167.7; 362.9
					157.9; 355.5

(Fig. 1i)						
87	4.29	5.71	30.00	60.00	77.9; 166.4; 364.3	H
					160.2; 358.8	C
89	2.14	2.86	15.00	80.00	78.4; 166.5; 368.7; 384.1	H
					161.7; 377.4	C

* mean of two cycles.

TABLE 2b

The results of DTA experiments for quaternary Bi-In-Sn-Zn alloys (samples 37, 38 and 68) acquired by DTA/DSC Pegasus 404 F1 (Netzsch, Germany)

Sample no.	Bi at. %	In at. %	Sn at. %	Zn at. %	Thermal arrests*, °C
(in Fig. 1d)					
Heating (H) and cooling (C) – rate 5 deg/min					
37	10.00	20.00	10.00	60.00	60.5; 63.1; 68.6; 76; 385.1; 420.0
					60.5; 64.2; 80.1; 385; 419.6
38	7.50	15.00	7.50	70.00	60.7; 70.8; 76.7; 384.3; 429.5
					64.6; 77.7; 384.4; 438.1
Heating (H) and cooling (C) – rate 2 deg/min					
37	10.00	20.00	10.00	60.00	58.8; 61.8; 67.0; 70.4; 80.8; 380; 385.1; 391.7; 431.2
					54.8; 57.2; 59.6; 77.4; 384.1; 436.1
38	7.50	15.00	7.50	70.00	58.7; 60.9; 67.2; 71.4; 79.5; 383.5; 385.5; 414.6
					54.7; 59.5; 78.1; 381.5; 383.2; 413.9
(in Fig. 1g)					
Heating (H) and cooling (C) – rate 5 deg/min					
68	9.64	12.86	7.50	70.00	63.3; 69.6; 79.9; 87.4; 386.9; 454.5
					60.9; 63.2; 66.7; 77.8; 389.7; 454.8
Heating (H) and cooling (C) – rate 2 deg/min					
68	9.64	12.86	7.50	70.00	59.0; 62.3; 72.4(b); 77.2; 81.9; 382.6; 389.8; 452.7
					55.0; 59.2; 65.2; 75.0; 385.6; 389.0; 453.5

* mean of two cycles, (b) – beginning of signal change.

3. Phase equilibria calculations

3.1. Thermodynamic modelling

The phases of the terminal solutions: ((Bi), (In), (Sn), (Zn)), α_1 , γ -InSn and Liquid (L) (denoted according to Table 3) are modelled with the use of a random substitutional solution. The Gibbs energy of these phases in the quaternary Bi-In-Sn-Zn system can be described in the following form:

$$G^\varphi = \varphi G^{ref} + \varphi G^{id} + \varphi G_{bin}^{ex} + \varphi G_{ter}^{ex}$$

$$(\varphi = (\text{Bi}), (\text{In}), (\text{Sn}), (\text{Zn}), \alpha_1, \gamma\text{-InSn and L}) \quad (1)$$

where: ${}^\varphi G^{ref} + {}^\varphi G^{id}$ is the sum of the mechanical mixture and the ideal mixing Gibbs energy terms, ${}^\varphi G_{bin}^{ex}$ is the excess Gibbs free energy of the binary alloys approximated from the Gibbs energies of the subsystems with the use of the Muggianu extrapolation formula [11] and ${}^\varphi G_{ter}^{ex}$ is the correction function of the ternary interactions in the alloys. The excess terms represent the composition dependence of the Gibbs energy of the phases (φ).

TABLE 3

Detailed information about the phases in the Bi-In-Sn-Zn system

Name, (symbol in phase diagram and in TDB database)	Pearson's symbol	Constitution
Liquid, L		(Bi, In, Sn, Zn)
Rhombohedral-A7, (Bi)	hR2	(Bi, In, Sn, Zn)
Tetragonal-A6, (In)	tI2	(Bi, In, Sn, Zn)
Bct-A5, (Sn)	tI4	(Bi, In, Sn, Zn)
Hcp-Zn (Zn)	hP2	(Bi, In, Sn, Zn)
BiIn	tP4	(Bi) _{1/2} (In) _{1/2}
Bi ₃ In ₅	tI32	(Bi) _{3/8} (In) _{5/8}
BiIn ₂	hP6	(Bi) _{1/3} (In) _{2/3}
Tetragonal-A6, (α_1)	tI2	(Bi, In, Sn)
γ -InSn, (γ)	hP5	(In, Sn)

The temperature dependence of the molar Gibbs energy of the pure elements, referred to the standard state, is given by the following expression, according to the Scientific Group Thermodata Europe (SGTE) [12].

The calculations were performed with the use of the following relations describing the excess Gibbs free energy of the quaternary alloys:

$${}^\varphi G_{bin}^{ex} = \sum_{i < j} x_i x_j {}^\varphi L_{ij} \quad (2)$$

representing Redlich-Kister sum of binaries [13], where:

$${}^\varphi L_{ij} = \sum_{k=0}^{m_{ij}} {}^\varphi L_{ij}^{(k)} (x_i - x_j)^k \quad (3)$$

and $x_{i,(j,k,l)}$ are the mole fractions of components i, j, k, l and L_{ij} are the interaction parameters of the binary alloys calculated in the course of optimization of the thermodynamic properties of the binary alloys. T is the temperature at the absolute scale. The coefficients ${}^\varphi L_{ij}^{(k)}$ of the Redlich-Kister equation [12] can also be a function of temperature, according to Eq. 4, and they were assessed on the basis of the available equilibrium experimental data.

$${}^\varphi L_{ij}^{(k)} = {}^\varphi A_{ij}^{(k)} + {}^\varphi B_{ij}^{(k)} T \quad (4)$$

In the case of the ternary alloys, the thermodynamic model was completed by the ternary interaction term:

$${}^\varphi G_{ter}^{ex} = \sum_{i < j < k} x_i x_j x_k {}^\varphi L_{ijk} \quad (5)$$

The composition dependence of the ternary interaction parameter ${}^\varphi L_{ijk}$ is given by the following assessed parameters ${}^\varphi L_{i,(j,k)}$:

$${}^\varphi L_{ijk} = x_i {}^\varphi L_i + x_j {}^\varphi L_j + x_k {}^\varphi L_k \quad (6)$$

In present calculations the quaternary interaction term was not taken into account.

The stoichiometric phases of the binary BiIn, Bi₃In₅ and BiIn₂ were described with the use of the simple assumption of the Kopp-Neumann rule [14], i.e. by a floating reference state:

$${}^0 G_{BiIn_q}^{Bi_p In_q} = p {}^0 G_{Bi}^{Rhombohedral-A7}(T) + q {}^0 G_{In}^{Tetragonal-A6}(T) + A_{pq} + B_{pq} \cdot T \quad (7)$$

where: p and q are the stoichiometric coefficients.

${}^0 G_{Bi}^{Rhombohedral-A7}(T)$ and ${}^0 G_{In}^{Tetragonal-A6}(T)$ are the standard SER functions for elemental Bi and In, respectively. The assessed enthalpy and entropy parameters A_{pq} and B_{pq} of the Gibbs free energy of the formation of the BiIn, Bi₃In₅ and BiIn₂ compounds taken from the assessment of Boa and Ansara [15] has been stored in the COST 531 ver. 3.0 database [16,17] of June 2008 version.

The thermodynamic description of quaternary Bi-In-Sn-Zn system was combined from assessments of six constituent ternary systems. It is a common practice for multicomponent systems because of lack of thermodynamic data. Detailed information on the phases in the Bi-In-Sn-Zn system is given in Table 3. The additional stable ternary and quaternary phases in the Bi-In-Sn-Zn system were not reported. The tin diamond crystal structure, which is stable below 286 K (13°C), is not considered here.

Phase equilibria in the Bi-In-Sn-Zn system were determined by the Pandat software ver. 8.1 (CompuTherm LLT, USA) [18, 19]. The parameters of the Gibbs energy of metals and the interaction parameters of the liquid and solid phases were taken from the COST531 ver. 3.1 data base [7]. In this database the assessment of the Bi-Zn system was based on Malakhov's system optimization [20] with an improvement of the new data for the solubility of Bi in the hcp-Zn terminal phase [21]. The thermodynamic modelling of Lee et al. [22], with some amendments of the In-Zn interaction parameters of the binary solid phases, was adopted for the In-Zn system. According to the assessment of Vizdal et al. [21], the datasets for the Bi-Sn and Bi-Sn-Zn systems were included in the COST531 database [7].

The interaction parameters for the liquid Bi-In-Zn solutions were elaborated on the basis of the experimental data which have been published in the recent years and the DTA data presented by Onderka et al. [4]. In the case of the Bi-In-Sn system, the thermodynamic parameters were taken from the assessment of Yoon et al. [23], with some later modification [7].

The ternary parameters of the liquid solution in the In-Sn-Zn system were obtained from the thermodynamic assessment of Cui et al. [24].

The parameters of Bi-In-Sn [8, 9] assessment was not used in present calculations because it contains new Bi-In assessment which has not been implemented in COST 531 ver. 3.0 database [16,17] up till the conclusion of present paper.

All the thermodynamic parameters which have been used in this work are listed in Table 4. Some positive binary interaction

TDB Database of Bi-In-Sn-Zn system (298.15-2000 K). Model parameters from COST 531 v.3.0 database [16, 17] and from [3].
The Gibbs free energy of pure elements (also SER) in phases were not enclosed

(Sn) (Bi, In, Sn, Zn)		${}^0G_{\text{Bi,Sn}}^{\text{Liquid}} = 500+1.5\cdot T$ [21]
${}^0G_{\text{Bi,In}}^{(\text{Sn})} = 5000$ [16]		${}^1G_{\text{Bi,Sn}}^{\text{Liquid}} = -100-0.135\cdot T$ [21]
${}^0G_{\text{Bi,Sn}}^{(\text{Sn})} = 3500-1.038\cdot T$ [21]		${}^0G_{\text{Bi,Zn}}^{\text{Liquid}} = 18265.1-8.6763\cdot T$ [20]
${}^1G_{\text{Bi,Sn}}^{(\text{Sn})} = -3710$ [21]		${}^1G_{\text{Bi,Zn}}^{\text{Liquid}} = -6061.2+0.7958\cdot T$ [20]
${}^0G_{\text{Bi,Zn}}^{(\text{Sn})} = 50\cdot T^{(a)}$		${}^2G_{\text{Bi,Zn}}^{\text{Liquid}} = -6422.6+11.7197\cdot T$ [20]
${}^0G_{\text{In,Zn}}^{(\text{Sn})} = 3500$ [16]		${}^3G_{\text{Bi,Zn}}^{\text{Liquid}} = 7227.4-9.2905\cdot T$ [20]
${}^0G_{\text{Sn,Zn}}^{(\text{Sn})} = 6514.8+25.7096\cdot T$ [28]		${}^4G_{\text{Bi,Zn}}^{\text{Liquid}} = 21123.1-27.1471\cdot T$ [20]
${}^0G_{\text{Bi,In,Sn}}^{(\text{Sn})} = -1000$ [23]		${}^5G_{\text{Bi,Zn}}^{\text{Liquid}} = -20747.6+22.0176\cdot T$ [20]
${}^1G_{\text{Bi,In,Sn}}^{(\text{Sn})} = -1000$ [23]		${}^6G_{\text{Bi,Zn}}^{\text{Liquid}} = -7600.4+13.1596\cdot T$ [20]
${}^2G_{\text{Bi,In,Sn}}^{(\text{Sn})} = -33000$ [23]		${}^0G_{\text{In,Sn}}^{\text{Liquid}} = -828.5+0.7602\cdot T-0.1212\cdot T\cdot \ln(T)$ [16]
Bi₃In₅ (Bi)_{0.375}(In)_{0.625}		${}^1G_{\text{In,Sn}}^{\text{Liquid}} = -115.6-1.40\cdot T$ [16]
${}^0G_{\text{Sn,Zn}}^{\text{Bi}_3\text{In}_5} = -544-4.12287\cdot T+0.375\cdot {}^0G_{\text{Bi}}^{\text{Rhomb}_A7} +$ $+ 0.625\cdot {}^0G_{\text{In}}^{\text{Tetrag}_A6}$ [15]		${}^0G_{\text{In,Zn}}^{\text{Liquid}} = 12401-4.4498\cdot T$ [22]
BiIn (Bi)_{0.5}(In)_{0.5}		${}^1G_{\text{In,Zn}}^{\text{Liquid}} = -3186+1.756\cdot T$ [22]
${}^0G_{\text{Bi,In}}^{\text{BiIn}} = -732.2-3.7906\cdot T+0.5\cdot {}^0G_{\text{Bi}}^{\text{Rhomb}_A7} + 0.5\cdot {}^0G_{\text{In}}^{\text{Tetrag}_A6}$ [15]		${}^2G_{\text{In,Zn}}^{\text{Liquid}} = 679$ [22]
BiIn₂ (Bi)_{0.3333}(In)_{0.6667}		${}^0G_{\text{Sn,Zn}}^{\text{Liquid}} = 19314.6-75.8995\cdot T+8.7514\cdot T\cdot \ln(T)$ [28]
${}^0G_{\text{Bi,In}}^{\text{BiIn}_2} = -481.1-4.188\cdot T+0.3333\cdot {}^0G_{\text{Bi}}^{\text{Rhomb}_A7} +$ $+ 0.6667\cdot {}^0G_{\text{In}}^{\text{Tetrag}_A6}$ [15]		${}^1G_{\text{Sn,Zn}}^{\text{Liquid}} = -5696.3+4.202\cdot T$ [28]
(Zn) (Bi, In, Sn, Zn)		${}^2G_{\text{Sn,Zn}}^{\text{Liquid}} = 1037.2+0.9836\cdot T$ [28]
${}^0G_{\text{Bi}}^{(\text{Zn})} = 9901-11.8\cdot T+ {}^0G_{\text{Bi}}^{\text{Rhomb}_A7}$ [3]		${}^0G_{\text{Bi,In,Sn}}^{\text{Liquid}} = 8743$ [23]
${}^0G_{\text{Bi,Sn}}^{(\text{Zn})} = 2000$ [16]		${}^1G_{\text{Bi,In,Sn}}^{\text{Liquid}} = 8743$ [23]
${}^0G_{\text{Bi,Zn}}^{(\text{Zn})} = 25000$ [16]		${}^2G_{\text{Bi,In,Sn}}^{\text{Liquid}} = 8743$ [23]
${}^0G_{\text{In,Sn}}^{(\text{Zn})} = 5000+2\cdot T$ [16]		${}^0G_{\text{Bi,In,Zn}}^{\text{Liquid}} = -11397.1+0.7976\cdot T$ [4]
${}^0G_{\text{In,Zn}}^{(\text{Zn})} = 23114$ [22]		${}^1G_{\text{Bi,In,Zn}}^{\text{Liquid}} = 7495.8$ [4]
${}^0G_{\text{Sn,Zn}}^{(\text{Zn})} = 33433.9-11.145\cdot T$ [16]		${}^2G_{\text{Bi,In,Zn}}^{\text{Liquid}} = -19544.9+16.22\cdot T$ [4]
γ-InSn (In, Sn)		${}^0G_{\text{Bi,Sn,Zn}}^{\text{Liquid}} = -17690.6+33.0\cdot T$ [21]
${}^0G_{\text{In}}^{\gamma\text{-InSn}} = 6000-2\cdot T+ {}^0G_{\text{In}}^{\text{Tetrag}_A6}$ [22]		${}^1G_{\text{Bi,Sn,Zn}}^{\text{Liquid}} = -2737.2-13\cdot T$ [21]
${}^0G_{\text{Sn}}^{\gamma\text{-InSn}} = 625.5-1.1350\cdot T+ {}^0G_{\text{Sn}}^{\text{Bct}_A5}$ [22]		${}^2G_{\text{Bi,Sn,Zn}}^{\text{Liquid}} = -19259.1+0.5\cdot T$ [21]
${}^0G_{\text{In,Sn}}^{\gamma\text{-InSn}} = -4916+1.6387\cdot T$ [22]		${}^0G_{\text{In,Sn,Zn}}^{\text{Liquid}} = -2001-3.61\cdot T$ [24]
LIQUID (Bi, In, Sn, Zn)		${}^1G_{\text{In,Sn,Zn}}^{\text{Liquid}} = 832-4.2\cdot T$ [24]
${}^0G_{\text{Bi,In}}^{\text{Liquid}} = -7165-0.3754\cdot T$ [15]		${}^2G_{\text{In,Sn,Zn}}^{\text{Liquid}} = -16257+21.2\cdot T$ [24]
${}^1G_{\text{Bi,In}}^{\text{Liquid}} = 1503.8-0.5418\cdot T$ [15]		(Bi) (Bi, In, Sn, Zn)
${}^2G_{\text{Bi,In}}^{\text{Liquid}} = 1221.2-1.6597\cdot T$ [15]		${}^0G_{\text{Bi,In}}^{(\text{Bi})} = 22500$ [15]
${}^3G_{\text{Bi,In}}^{\text{Liquid}} = -1627+2.764\cdot T$ [15]		${}^0G_{\text{Bi,In}}^{(\text{Bi})} = 19720-22.6\cdot T$ [21]
		${}^1G_{\text{Bi,In}}^{(\text{Bi})} = -5760+13.834\cdot T$ [21]
		${}^0G_{\text{Bi,Zn}}^{(\text{Bi})} = 10000$ [16]
		${}^0G_{\text{In,Zn}}^{(\text{Bi})} = 3500$ [16]

${}^0G_{Sn,Zn}^{(Bi)} = 33433.9 - 11.1447 \cdot T$	[16]	${}^1G_{Bi,In,Sn}^{(In)} = 8000$	[23]
${}^0G_{Bi,In,Zn}^{Liquid} = 387000$	[21]	$\alpha_1 (Bi, In, Sn)$	
(In) (Bi, In, Sn, Zn)		${}^0G_{Bi}^{\alpha_1} = {}^0G_{Bi}^{Rhomb_{A7}} + 49.93$	[16]
${}^0G_{Bi,In}^{(In)} = 5646.8 - 26.868 \cdot T$	[15]	${}^0G_{In}^{\alpha_1} = 123 - 0.1988 \cdot T + {}^0G_{In}^{Tetrag_{A6}}$	[15]
${}^0G_{Bi,Sn}^{(In)} = 4000$	[16]	${}^0G_{Sn}^{\alpha_1} = 5510 - 8.46 \cdot T + {}^0G_{Sn}^{Bct_{A5}}$	[16]
${}^0G_{Bi,Zn}^{(In)} = 10000$	[16]	${}^0G_{Bi,In}^{\alpha_1} = 3308.4 - 27.257 \cdot T$	[16]
${}^0G_{In,Sn}^{(In)} = 2036 - 7.88412 \cdot T$	[16]	${}^1G_{Bi,In}^{\alpha_1} = -1661.4 - 1.644 \cdot T$	[16]
${}^0G_{In,Zn}^{(In)} = 4430 - 4.4498 \cdot T$	[22]	${}^0G_{Bi,Sn}^{\alpha_1} = 2000$	[16]
${}^1G_{In,Zn}^{(In)} = 9717$	[22]	${}^0G_{In,Sn}^{\alpha_1} = 950.4 - 6.1461 \cdot T$	[16]
${}^0G_{Sn,Zn}^{(In)} = 33433.9 - 11.1447 \cdot T$	[16]	${}^0G_{Bi,In,Sn}^{\alpha_1} = 13000$	[23]
${}^0G_{Bi,In,Sn}^{(In)} = 10000$	[23]	${}^1G_{Bi,In,Sn}^{\alpha_1} = 21000$	[23]

(a) parameters introduced to restrict the stability range of certain solid phases. All parameters not defined in model description are equal zero

parameters have been introduced to restrict the stability range of certain solid phases. Such parameters are marked (a) in Table 4.

The results of the invariant reaction calculations of the quaternary Bi-In-Sn-Zn system are presented in Table 5.

TABLE 5

The invariant reactions of the quaternary Bi-In-Sn-Zn system calculated in present work

Reaction type	Reaction	T, °C	Composition of liquid phase			
			at.% Bi	at.% In	at.% Sn	at.% Zn
U ₁	L+(Bi)↔(Sn)+(Zn)+BiIn	79.3	38.4	39.5	21.17	0.93
U ₂	L+(Sn)↔(Zn)+γ-InSn+BiIn	65.8	25.1	53.3	20.95	0.65
U ₃	L+BiIn↔(Zn)+γ-InSn+Bi ₃ In ₅	56.4	21.6	58.4	19.48	0.52
U ₄	L+Bi ₃ In ₅ ↔(Zn)+γ-InSn+BiIn ₂	54.8	20.3	59.5	16.7	0.50
E	L↔(Zn)+α ₁ +γ-InSn+BiIn ₂	54.2	19.5	60.1	19.9	0.50

(U₁ – pseudoperitectic reaction, E – quaternary eutectic reaction [3]).

4. Discussion

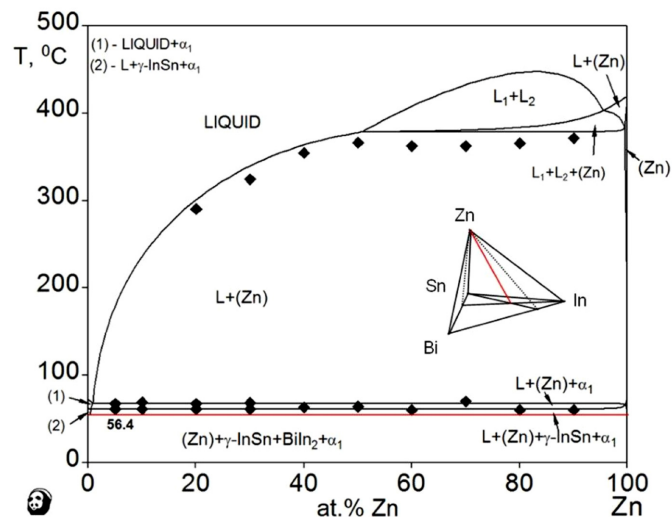
The results of the comparison of the calculated isopleths with the experimental data obtained in the presented work are shown in Figs 1a-i. The calculated sections of the Bi-In-Sn-Zn system agree very well with the experimental data obtained from two different thermal analyzers. The slight differences between calculated and experimentally acquired invariant reactions ($\pm 7^\circ\text{C}$) were observed (Fig. 1c).

The experimental data of Sabbar et al. [25] obtained for the In-Sn-Zn system suggest solubility of Zn in the γ-InSn phase. Unfortunately, the amount of data is too limited to model this

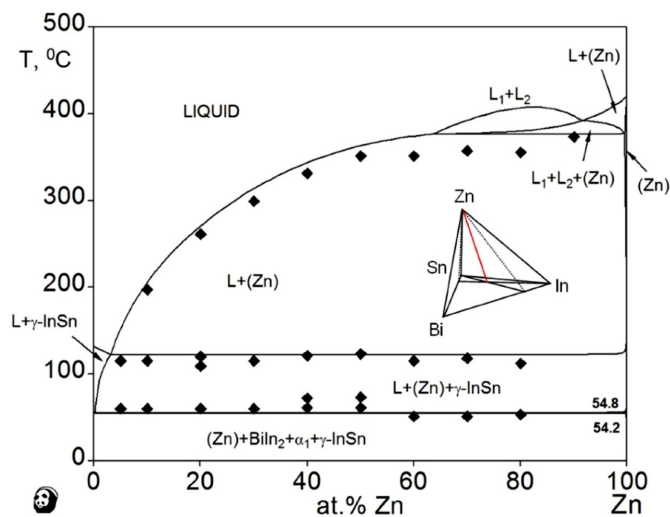
solubility. Additionally, to describe the Zn solubility in the γ-InSn phase the unknown expression for ${}^0G_{Zn}^{\gamma\text{-InSn}}$ should be estimated. So, such Zn solubility can expand stability of γ-InSn into compositions of ternary system. Also, the solubility of Zn in α₁ was not modeled because of the unknown solubility data. The unknown expression for ${}^0G_{Zn}^{\alpha_1}$ can be estimated from ${}^0G_{Zn}^{Hcp_{A3}}$. Generally, the mutual solubility of tin and zinc is very low and cannot cause any significant changes in the invariant reactions.

Another problem is an adequate description of some other ternary systems. For example, it was not possible to model the correct temperature dependence of the mixing enthalpy of the liquid phase in the In-Sn-Zn system. In some regions of the diagram, the calculated temperature dependence is opposite to that which is experimentally observed [26]. In these regions, the experimentally determined mixing enthalpy increases with the increasing temperature, whereas the calculated mixing enthalpy decreases with the increasing temperature for all concentrations of the system. This negative temperature dependence in the calculated mixing enthalpy originates from the term $T \cdot \ln(T)$ introduced in the thermodynamic description of the Sn-Zn system, but in the ternary system, apparently the temperature dependence shifts. To solve this problem, Moleans et al. [3] proposed to include the $T \cdot \ln T$ term in the description of the liquid ternary solution. However, the experimental data available up till now covers too narrow temperature and composition ranges, and their accuracy seems to be not sufficient, to do it.

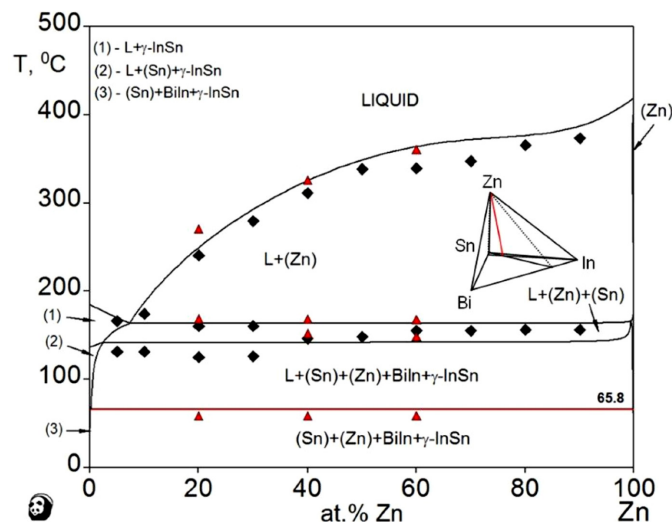
The phase diagram of the Bi-In-Sn ternary system was re-examined in the study of Yoon et al. [23] and the calculated phase diagram has both similarities and dissimilarities to the previously determined one. While the phase-transition temperatures are generally in agreement, the new phase diagram represents a different nature of the quasi-binaries, the number and type of the invariant reactions, and the reaction paths [23]. It is found that the Sn-BiIn and Sn-BiIn₂ quasi-binary systems are not pseudo-binaries, which eventually leads to a phase diagram with only



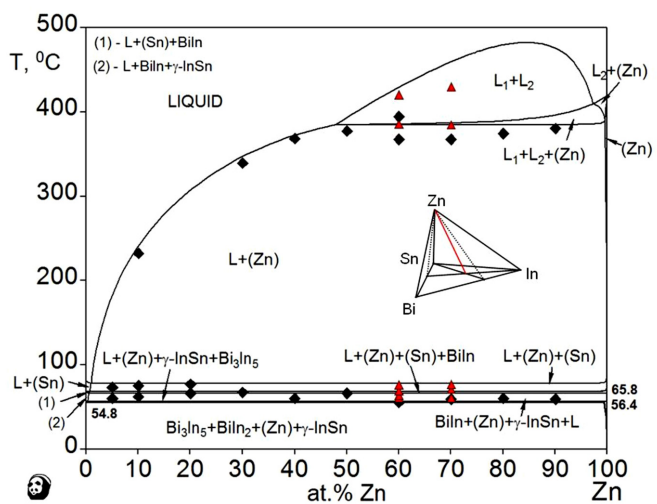
a)



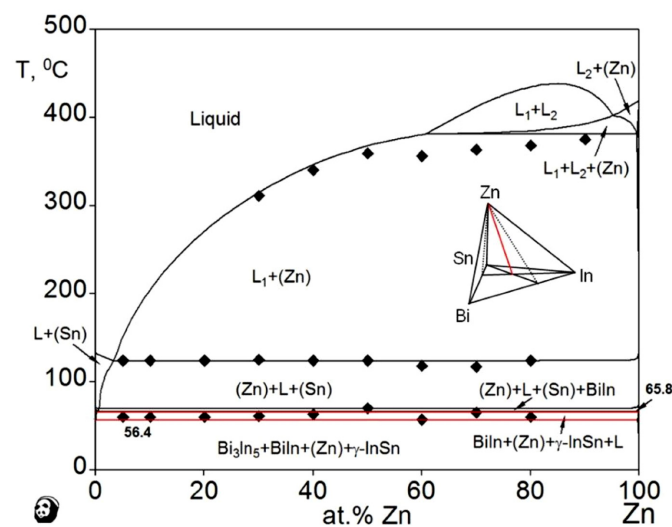
b)



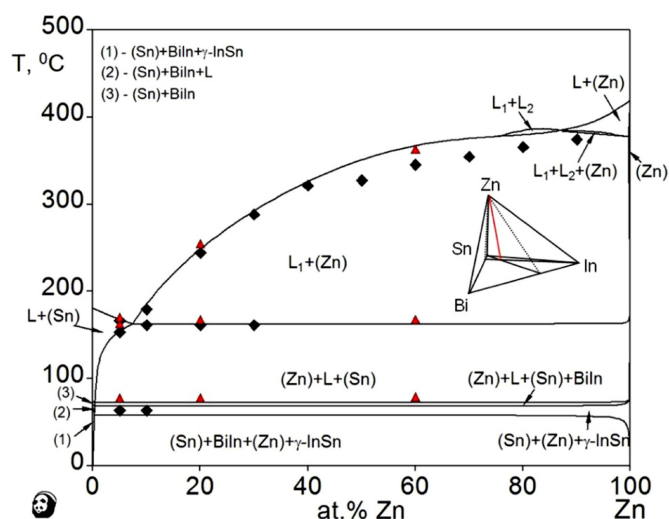
c)



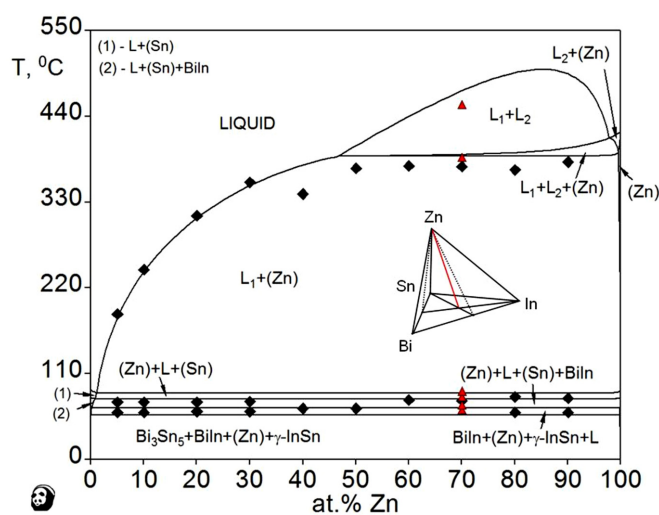
d)



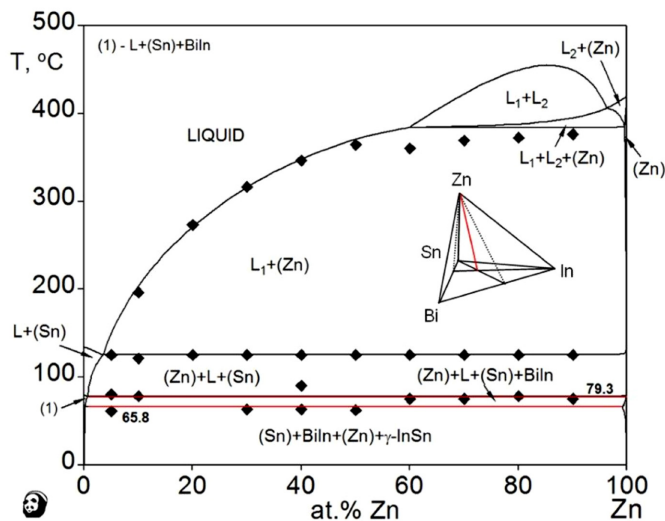
e)



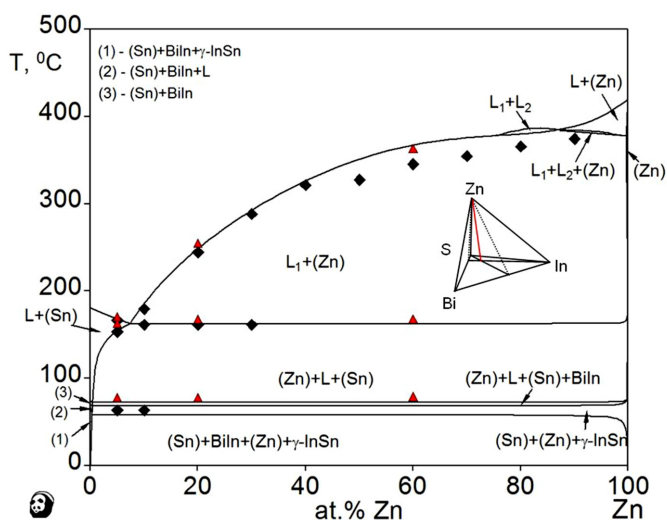
f)



g)



h)



i)

- Bi:In = 0.25 and Bi:Sn = 0.6.
- $X_{\text{Bi}}:X_{\text{In}} = 0.25$ and $X_{\text{Bi}}:X_{\text{Sn}} = 0.20$. The quaternary eutectic temperature E equals 54.2°C (327.3 K).
- $X_{\text{Bi}}:X_{\text{In}} = 0.25$ and $X_{\text{Bi}}:X_{\text{Sn}} = 0.067$. (Additionally, Table 2a).
- $X_{\text{Bi}}:X_{\text{In}} = 0.5$ and $X_{\text{Bi}}:X_{\text{Sn}} = 1$. (Additionally, Table 2b).
- $X_{\text{Bi}}:X_{\text{In}} = 0.50$ and $X_{\text{Bi}}:X_{\text{Sn}} = 0.33$.
- $X_{\text{Bi}}:X_{\text{In}} = 0.5$ and $X_{\text{Bi}}:X_{\text{Sn}} = 0.11$. (Additionally, Table 2a).
- $X_{\text{Bi}}:X_{\text{In}} = 0.75$ and $X_{\text{Bi}}:X_{\text{Sn}} = 1.286$. (Additionally, Table 2b).
- $X_{\text{Bi}}:X_{\text{In}} = 0.75$ or $X_{\text{Bi}}:X_{\text{Sn}} = 0.43$.
- $X_{\text{Bi}}:X_{\text{In}} = 0.75$ and $X_{\text{Bi}}:X_{\text{Sn}} = 0.143$. (Additionally, Table 2a)

Fig. 1. Comparison of DTA data results (Tables 1a-i and Tables 2a-b) with the phase equilibria lines calculated of Bi-In-Sn-Zn system for the mole fraction ratios (♦ – data acquired by Q-1500 D, ▲, △ – heating data acquired by Netzsch 404 DTA (5 deg/min and 2 deg/min, respectively are not recognizable in the diagram scale)

one eutectic reaction. Other invariant reactions (eight in total) are determined to be peritectic and peritectoid [23].

In Figs 2a-b, the sequences of five thermal arrests obtained from the 2 deg/min DTA measurements of samples no. 37 and 38 are shown. Such the sequence is consistent with the calculated invariant and univariant lines of diagram in Fig. 1d, within the range of $\pm 3^\circ\text{C}$. It should be noted that difference between calculated two pairs of invariant and univariant lines is less than 3°C . So the match is very good. The resolution of Q-1500 D experimental data DTA data was enough (5 deg/min) to determine only one or two (5-20 at.% Zn) thermal arrests in this temperature range.

A similar low temperature sequence of four thermal arrests was observed on the DTA curve (Fig. 3) obtained for sample no. 68. Again, this sequence is well-matched with the calculated results of four univariant lines in temperature range $\sim 55\text{-}87^\circ\text{C}$ ($328\text{-}360\text{ K}$) shown in Fig. 1g.

The high temperature thermal arrests observed for samples 37 and 38 in both the heating and the cooling cycles (Fig. 1d) can be connected with the boundary of miscibility gap. It should be noted that taking into account thermodynamic description of quaternary alloys combined from assessments of constituent ternaries the consistence of calculated and experimental data are very good.

In all above described samples the difference between the calculated and experimental data of liquidus and the end of (Zn) melting acquired by Q-1500 D can be observed in the part of higher Zn contents. Also the calculated miscibility gap boundaries are higher (max. 25°C) than DTA results (Netzsch 404 DTA – Fig. 1d). The experimental error in this part of diagram can be connected with relatively high Zn partial pressure. It should be also noted that in present work it was possible to recognize in heating and cooling cycles the boundary of miscibility gap (Fig. 1d and 1g) which is known to be relatively difficult to detect.

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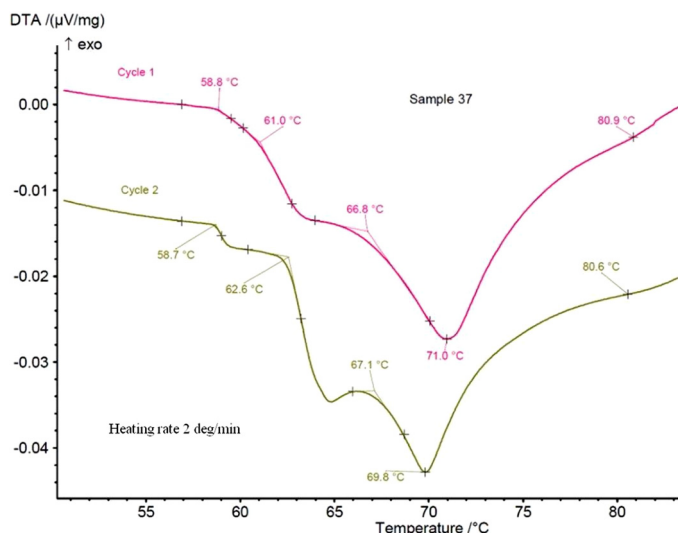


Fig. 2a. Low temperature part of DTA chart for two heating cycles of sample 37 ($X_{\text{Bi}}:X_{\text{In}} = 0.50$, $X_{\text{Bi}}:X_{\text{Sn}} = 1$ and $X_{\text{Zn}} = 0.60$)

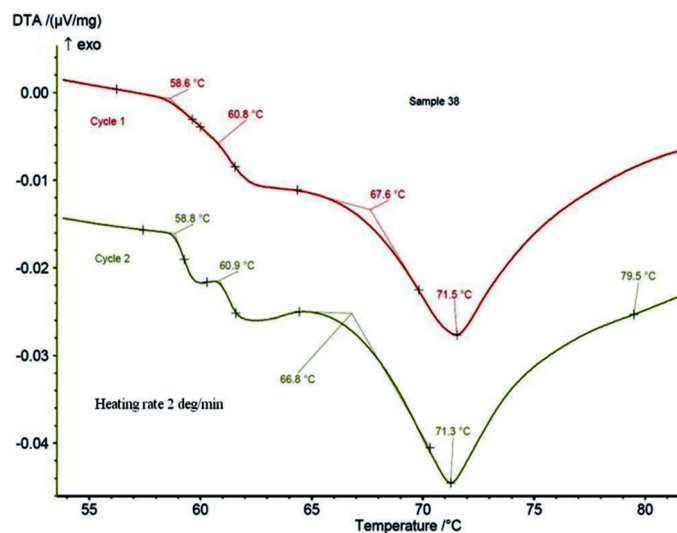


Fig. 2b. Low temperature part of DTA chart for two heating cycles of sample 38 ($X_{\text{Bi}}:X_{\text{In}} = 0.50$, $X_{\text{Bi}}:X_{\text{Sn}} = 1$ and $X_{\text{Zn}} = 0.70$)

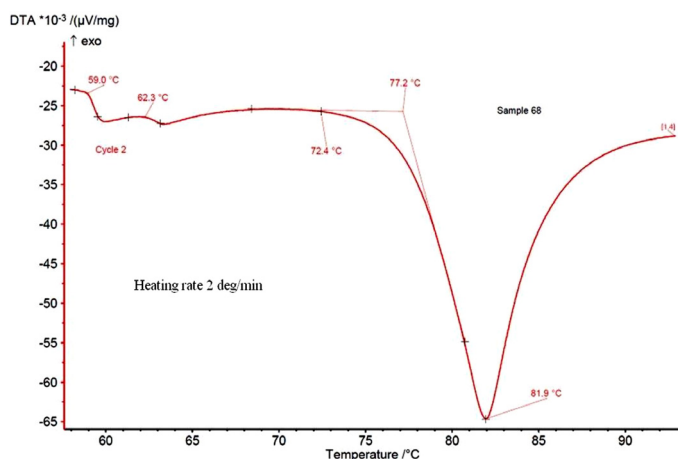


Fig. 3. Low temperature part of DTA chart for two heating cycles of sample 68 ($X_{\text{Bi}}:X_{\text{In}} = 0.75$, $X_{\text{Bi}}:X_{\text{Sn}} = 1.286$ and $X_{\text{Zn}} = 0.7$) – (Table 2b)

The comparison between the experimental data of Yoon et al. [2] and the calculated isopleths of Sn-Bi-2In-9Zn and Sn-Bi-5In-6Zn (in mas. %) is presented in Figs 5a-b, respectively. The composition dependence of the (Sn) dissolution limit is in very good agreement with the calculated results. Moreover, the difference in the experimental and the calculated temperatures of solidus and liquidus observed in Figs 4a-b can be connected with the unusual method of analysis of the thermal arrests found on the DSC curves performed by Yoon et al. [2].

Two years later Yoon et al. [23] reported the thermodynamic calculation of the Bi-In-Sn ternary system by the CALPHAD method despite the absence of experimental thermodynamic data available for the system. The only experimental data on this system were published by Witusiewicz et al. [8] and Rex et al. [9]. Their work gives a new data set for the ternary system

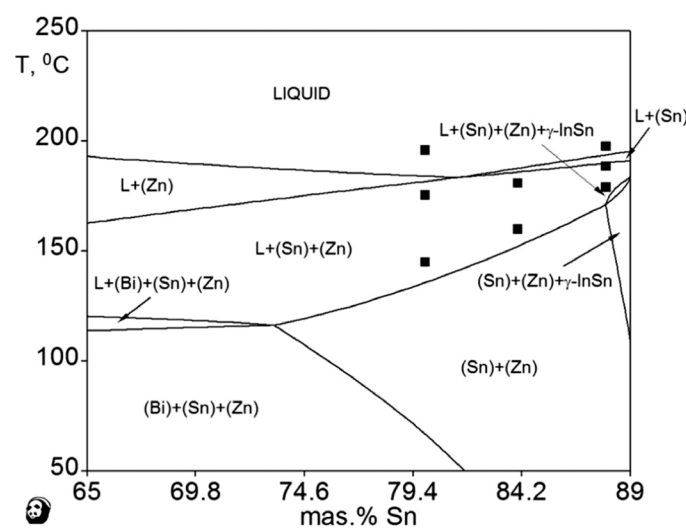
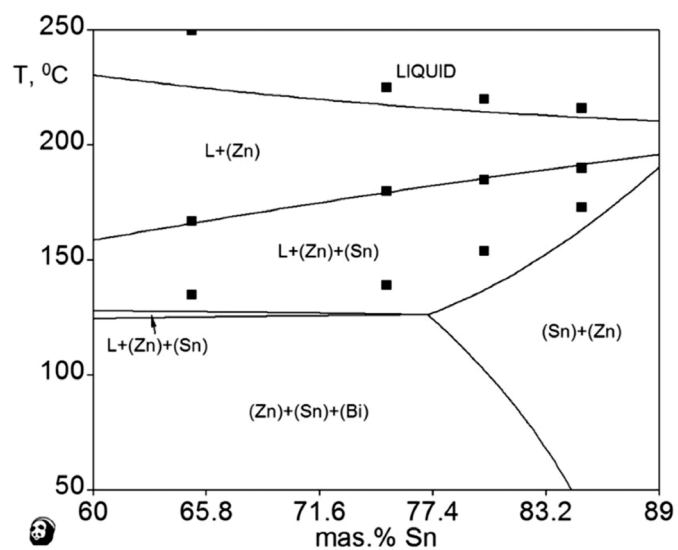


Fig. 4. Calculated isopleths of Bi-In-Sn-Zn system: (a) Sn-Bi-2In-9Zn and (b) Sn-Bi-5Sn-6Zn alloys compared with the temperature arrests measured by DSC method [2]

considering the β phase (BiIn_9), which is in better agreement with the experimental data than the earlier calculation. They also accepted the limiting solubility of Sn in the binary BiIn , Bi_3In_5 and BiIn_2 intermetallic phases. The solubility studies of Zn in the γ - InSn and α_1 phase together with the critical elaboration by Witusiewicz et al. [27] which allow, in the future, to update the description of the Bi-In-Sn-Zn system.

5. Conclusions

With the application of the thermal differential analysis (DTA), studies of the phase transition temperatures were performed for 90 of Bi-In-Sn-Zn alloys, for the mole fraction ratio of Bi:In equaling 0.25, 0.5 and 0.75 and for the ratio Bi:Sn equaling 0.067, 0.111, 0.143, 0.2, 0.333, 0.428, 0.5, 0.6, 1.286. The reliability of DTA data was tested by two types of thermal devices: Q-1500 D (Paulik-Paulik-Erdey) (MOM, Hungary) and DSC/DTA Pegasus 404 F1 (Netzsch, Germany).

Based on the critical thermodynamic evaluations (assessments) of the binary and ternary subsystems available in the literature, the calculations of the equilibrium lines of the phase transition were conducted in the quaternary Bi-In-Sn-Zn system, and compared with those measured in these studies. This is the standard practice to assess the multicomponent system from assessments of constituent ternary systems and to make some control experiments (also DTA) to check the accuracy of such description. The present work can be treated as such check of coherency of optimized ternary parameters in quaternary system database. And it seems that this check is positive because the comparative analysis of the experimental and the calculated data of the transition temperatures showed a good agreement of both sets of data.

The Bi-In-Sn-Zn alloys are important for lead-free soldering, and so the future studies should concentrate on the thermodynamic investigations of solid alloys, as their experimental data are very poor. Also, the determination of the regions of the phases' coexistence for the new elaboration of the Bi-In-Sn-Zn system is required.

Acknowledgement

This research has been supported by grant N N507 4572 37 from the Polish Ministry of Science and Higher Education.

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