

INVESTIGATIONS OF MASS TRANSFER IN ANNULAR GAS-LIQUID FLOW IN A MICROREACTOR

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Dedicated to Prof. Ryszard Pohorecki on the occasion of his 80th birthday

The paper presents an investigation of mass transfer in gas-liquid annular flow in a microreactor. The microreactor had a meandered shape with a square cross-section of the channel ($292 \times 292 \mu\text{m}$, hydraulic diameter $292 \mu\text{m}$) and 250 mm in length. The rate of CO_2 absorption from the CO_2/N_2 mixture in NaOH (0.1 M, 0.2 M, 0.7 M, 1.0 M and 1.5 M) water solutions was measured. Two velocities of gas flow and two velocities of liquid flow were used. In two cases a fully developed annular flow at the beginning of the channel was observed, whilst in two cases annular flow was formed only in about 2/3 of the microchannel length. Based on the measurements of CO_2 absorption rate, the values of volumetric liquid – side mass transfer coefficients with the chemical reaction were determined. Then physical values of coefficients were found. Obtained results were discussed and their values were compared with the values predicted by literature correlations.

Keywords: mass transfer, annular flow, microreactor

1. INTRODUCTION

Multiphase and homophase flows in microstructures play a significant role in many branches of contemporary technology. More and more frequently microreactors are used in producing a particular product, e.g. photochemical gas-liquid reactions (Ehrich et al., 2002), sulphonation and oxidation of hydrocarbons (Hessel et al., 2005), methane oxidation to formaldehyde (Zhang et al., 2011). Analytical methods (lab on a chip) constitute a great deal of microstructure usage (Churski et al., 2012), but also e.g. detection of gas solubility and diffusivity in liquids (Abolhasani et al., 2012; Sobieszuk and Pilarek, 2012). Due to the fact that high values of the interfacial area are observed in heterophase systems in microreactors with gas-liquid flow, they are widely examined as an innovative gas-liquid apparatus. There are a lot of published papers based of fundamental investigations of such a flow. In particular, it concerns hydrodynamics (Abadie et al., 2012; Pohorecki et al., 2008), pressure drop (Cygański et al., 2012; Saisorn and Wongwises, 2010), mixing (Commenge et al., 2006; Kolb et al., 2013; Malecha et al., 2009) and heat exchange (Hecht and Kraut, 2010). The greatest number of papers concerns mass transfer (Kashid et al., 2011; Sobieszuk et al., 2011; Yao et al., 2015). These investigations are conducted both for open and closed microchannels. In a closed channel in papers concerning hydrodynamics, all authors reported the presence of two main flow regimes, i.e. Taylor (slug) and annular ones. The Taylor flow is very stable and is present in a large range of media flow (Hessel et al., 2005). Therefore, it is the most interesting flow to carry out reactions in heterophase gas-liquid and liquid – liquid systems.

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The annular flow is definitely a more rarely examined regime. It is a consequence of a short time of phase contact. However, such a situation might be necessary and then the annular flow proves to be the appropriate one. A short time of the phase contact in the annular flow was used by Zhang et al. (2011) in investigating the kinetics of oxidation reaction of methane to formaldehyde. When enzymes were used in the liquid phase during the formation of slug flow, foaming was observed (preliminary experiments) making determination of the interfacial area impossible. During the creation of the annular flow, the foaming phenomenon was not observed; hence this regime might be dedicated to specific measurements of the rate of mass transfer in enzymatic reactions.

The present paper is focused on mass transfer in an annular flow, and more specifically, on determining the values of mass transfer coefficients. Furthermore, a microchannel of a meandered shape, and not the most typically described straight one, was used in experiments.

2. EXPERIMENTAL SETUP

The microreactor was made of glass and had a Y-shaped and meandered shape of its main channel. The cross-section of all the channels was square and hydraulic diameter was 0.292 mm. The main channel length was 250 mm. A scheme of meandered microreactor is shown in Fig 1a. A phase separator, in which liquid is separated from gas due to papillary rise, was applied. The separator scheme was presented in Fig 1b. The microreactor and the separator were fabricated by Mikrogas Chemtech GmbH from Mainz (Germany). The experimental set-up is shown in Fig. 2. The gases (nitrogen and carbon dioxide) were supplied from cylinders through reducing valves and gas flow regulators. Both gases were mixed before the inlet to the microchannel. Liquid was supplied by a syringe pump. The liquid flow was measured using a measuring cylinder and a stop watch. The two phase mixture leaving the microchannel was separated in a capillary separator. The outlet gas was directed to a gas chromatograph (Shimadzu GC-2014). The flow was recorded by a CMOS high speed camera (1200hs). The system allowed to record frames with the resolution of 1280×1024 pixels and the speed in the range from 500 to 4000 frames per second. The pixel size was about 10 μm. The microreactor was lit using a 500 W halogen lamp with a light diffuser. All the data were measured at room temperature. To eliminate possible channel heating, the halogen lamp was used only for several seconds. The rates of absorption of CO₂ from a CO₂/N₂ mixture into an aqueous solution of NaOH (0.1 M, 0.2 M, 0.7 M, 1.0 M and 1.5 M) were measured. Two values of the gas flow (6.0 ml/min, 6.5 ml/min) and two values of the liquid flow (0.12 ml/min, 0.17 ml/min) were used. In all conditions the rate of absorption was measured twice: once for the entire microreactor, and for the second time using only the phase contactor and the separator, without the main gas – liquid microchannel. This method was used by Yue et al. (2007), and also in our earlier papers (Sobieszuk et al. (2008) or Sobieszuk et al. (2011)). The CO₂ absorption rate was calculated from gas concentration changes between the inlet and the outlet compositions. The inlet gas concentration was about 38%.

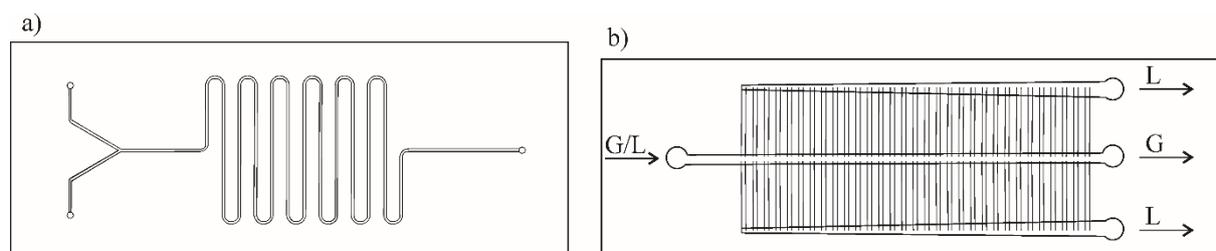


Fig. 1. Scheme of a) meandered microreactor, b) capillary phase separator

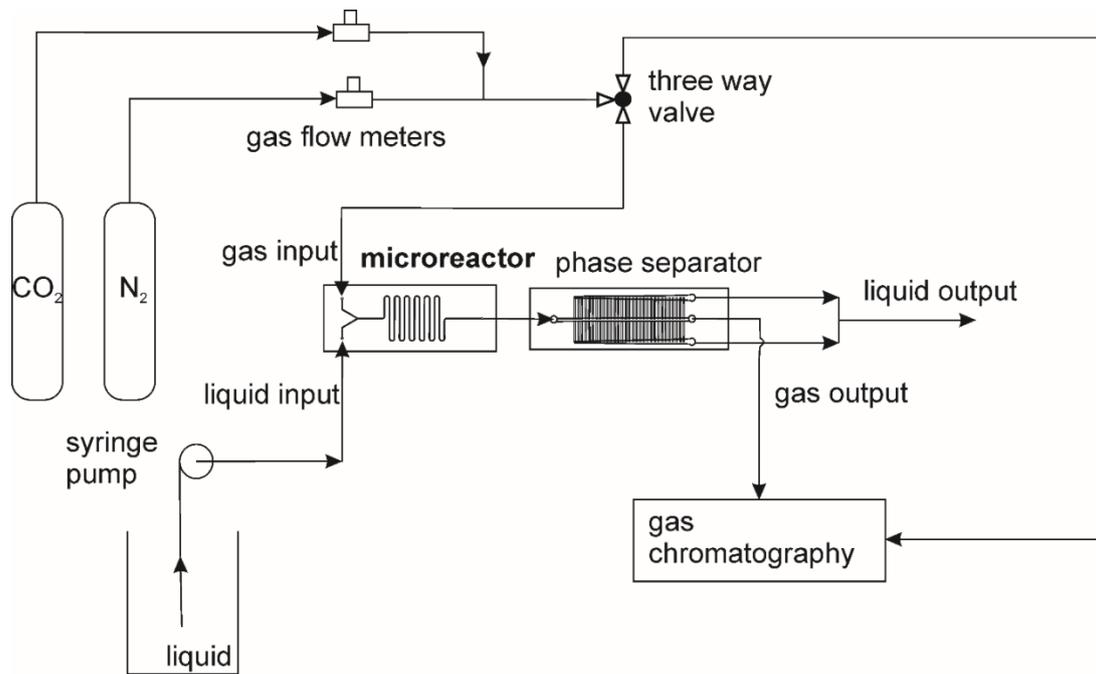


Fig. 2. Experimental set-up

3. METHODOLOGY

During the experiments the outlet CO₂ concentration was measured for a constant volumetric gas and liquid flow values, a constant CO₂ concentration at the gas inlet, and a constant NaOH concentration at the liquid inlet. Then the inlet (p_{A1}) and outlet (p_{A2}) CO₂ partial pressures were calculated. Finally, the rate of carbon dioxide absorption accompanied by a chemical reaction was found:

$$N_A^* = \frac{Q_{G1} p_{A1} - Q_{G2} p_{A2}}{RT} \quad (1)$$

Sobieszuk et al. (2014) report that for the Taylor flow in microchannels, gas side resistance of mass transfer (for CO₂) can be neglected. Therefore, the partial pressure of CO₂ at the gas inlet and outlet from the apparatus is equal to that at the interface at the inlet and outlet from the apparatus. Furthermore, the thermodynamic equilibrium described by the Henry's law was assumed at the interface:

$$\begin{aligned} c_{Ai1} &= H p_{Ai1} = H p_{A1} \\ c_{Ai2} &= H p_{Ai2} = H p_{A2} \end{aligned} \quad (2)$$

The absorption rate is defined by:

$$N_A^* = k_L^* a c_{Ai}^{AV} V \quad (3)$$

and c_{Ai}^{AV} is the logarithmic average of gas concentration at the interfacial area:

$$c_{Ai}^{AV} = \frac{c_{Ai1} - c_{Ai2}}{\ln \frac{c_{Ai1}}{c_{Ai2}}} \quad (4)$$

Consequently, from Eqs. (1) and (3) the experimental values of $k_L^* a$ can be estimated.

In order to determine physical values of the coefficients, it was assumed that the reaction is infinitely fast, i.e. the condition is fulfilled:

$$\frac{\sqrt{k_2 c_{B0} D_A}}{k_L} \gg \frac{c_{B0} D_B}{bc_{Ai}^{AV} D_A} \quad (5)$$

When the condition (5) is fulfilled, the enhancement factor is defined as follows:

$$E = 1 + \frac{c_{B0} D_B}{bc_{Ai}^{AV} D_A} \quad (6)$$

Since:

$$E = \frac{k_L^* a}{k_L a} \quad (7)$$

the values of the physical volumetric mass transfer coefficients were determined from Eq. (7). In order to verify the correctness of the assumption (5), k_L values need to be known. Therefore, the interfacial area during annular flow was first estimated. The cross-section of the flow in the square microchannel is presented in Fig. 3.

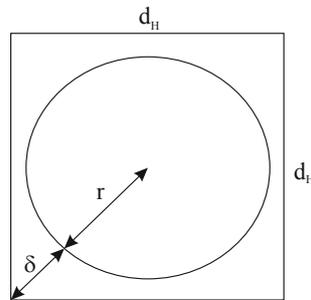


Fig. 3. Annular flow cross-section

Kreutzer et al. (2005) reported that in the rectangular channel, the thickness of the liquid film in microchannel corners can be determined from:

$$\frac{\delta}{d_H} = 0.35 - 0.25 \exp(-2.25 Ca^{0.445}) \quad (8)$$

Thus the gas stream is of the shape of a cylinder with the base radius equal to

$$r = 0.5 d_H \sqrt{2} - \delta \quad (9)$$

Finally, the interfacial area was determined as:

$$a = \frac{2\pi rL}{V} \quad (10)$$

where L is the microchannel length. Using the a value, the experimental values of k_L were calculated and it was confirmed that the condition (5) was always fulfilled.

The obtained values of the physical mass transfer coefficients for the annular flow were compared with those resulting from the Higbie's model and a correlation formulated by Yue et al. (2007).

According to Higbie:

$$k_L = 2\sqrt{\frac{D_A}{\pi\tau}} \quad (11)$$

τ is the time of phase contact, which can be defined in the microreactor as:

$$\tau = \frac{L}{u_G + u_L} \quad (12)$$

Yue et al. (2007) investigated mass transfer in a straight microchannel of a rectangular cross section (1 mm in depth, 0.5 mm in width). They suggested the following correlation for the slug-annular and churn flow:

$$Sh a d_H = 0.058 Re_G^{0.344} Re_L^{0.912} Sc^{0.5} \quad (13)$$

The physicochemical properties of CO₂-H₂O-NaOH system were widely published. From the available data the following were used:

- diffusivities D_A and D_B : Pohorecki and Moniuk (1988a), Danckwerts (1970),
- Henry's constant H : Barret (1966),
- reaction rate constant k_2 : Pohorecki and Moniuk (1988b),
- viscosity of NaOH water solutions μ_L : Pohorecki and Moniuk (1991),
- density of NaOH water solutions ρ_L : Pohorecki and Moniuk (1991).

4. RESULTS AND DISCUSSION

The investigation of absorption rate was carried out in the annular flow. It was determined that the fully developed flow regime was obtained at the gas flow rate of 6 ml/min. The experiments were marked as series 1, and the photograph of this flow was presented in Fig. 4.

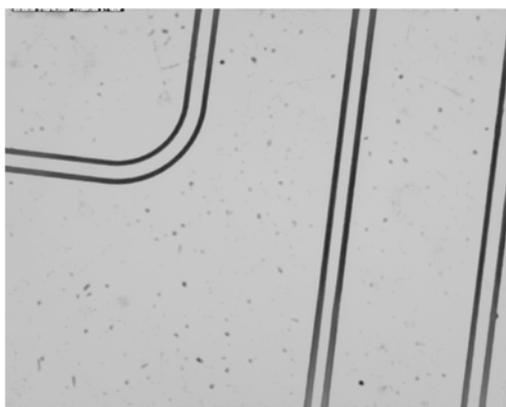


Fig. 4. Developed annular flow (series 1)

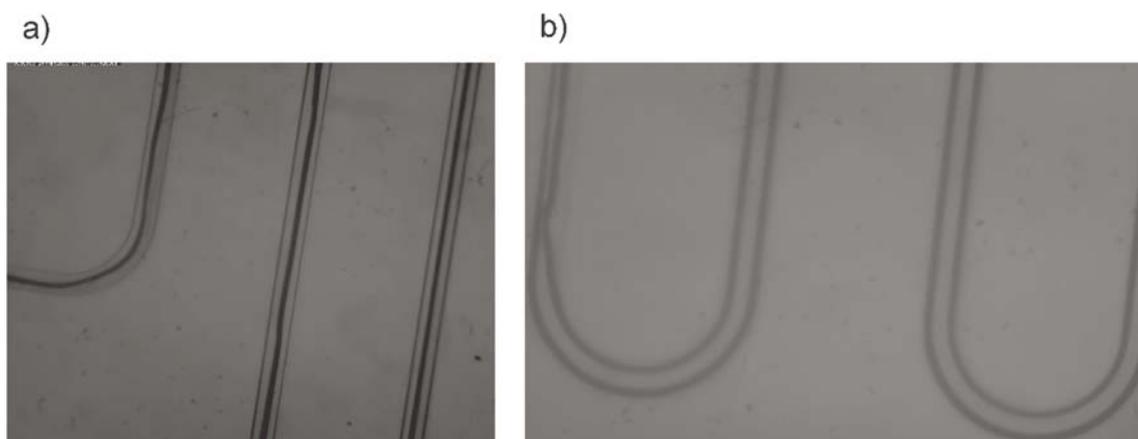


Fig. 5. Initially undeveloped annular flow a) and stability flow b) (series 2)

For the gas flow rate equal to 6.5 ml/min, the annular flow at the beginning of the microchannel was unstable (Fig. 5a); its stability was reached at about 1/3 of microchannel length (Fig. 5b). The experiments were marked as series 2.

The absorption rate was determined based on Eq. (1). A dependency of CO₂ absorption rate in the function of superficial velocity of two-phase flow was presented in Fig. 6.

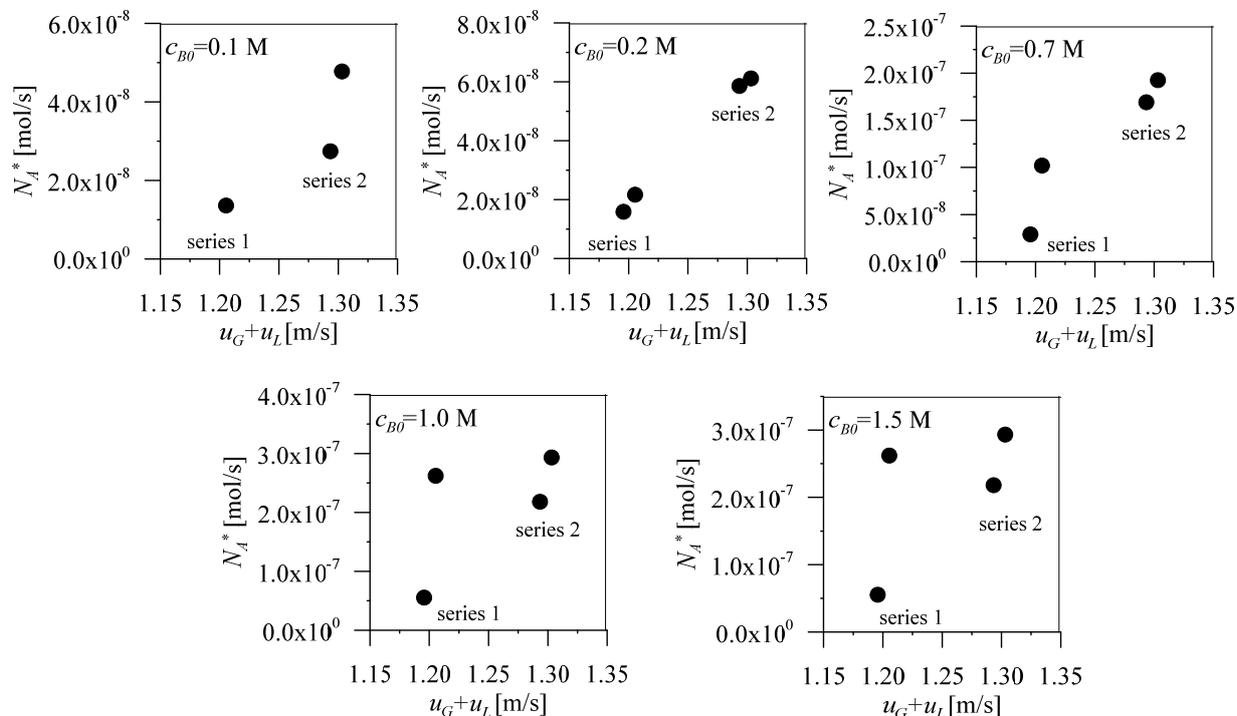


Fig. 6. Absorption rate in superficial velocity function of two-phase flow

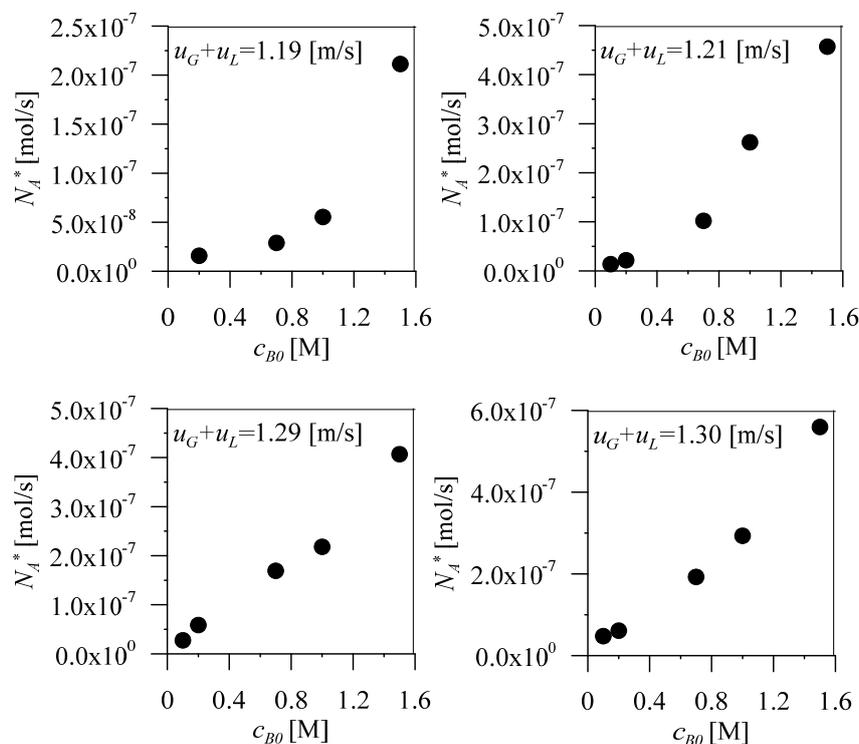


Fig. 7. NaOH concentration influence on the absorption rate

As show in Fig. 6 the absorption rate increases with the increasing media flow. At the higher NaOH (1.0 M and 1.5 M) concentration, an insignificant decrease in the absorption rate for series 2 was observed. It occurred because during incompletely developed annular flow, the interfacial area decreased. In Fig. 7, NaOH concentration influence on the absorption rate in the same hydrodynamic conditions was presented. Here it was proven that the NaOH concentration increase results in the increase of the absorption rate.

Values of the volumetric mass transfer coefficients obtained from Eq. (7) are presented in Fig. 8a. As can be seen, k_{La} values are relatively insignificant in comparison with the values observed in the Taylor flow (Sobieszuk et al., 2008; Sobieszuk et al., 2011; Yue et al., 2007). Some differences in the k_{La} values at the same superficial two-phase flow velocity, can be observed. They may also result from different physicochemical properties of NaOH solutions. The concentrations of NaOH were changed in a wide range (0.1-1.5 M).

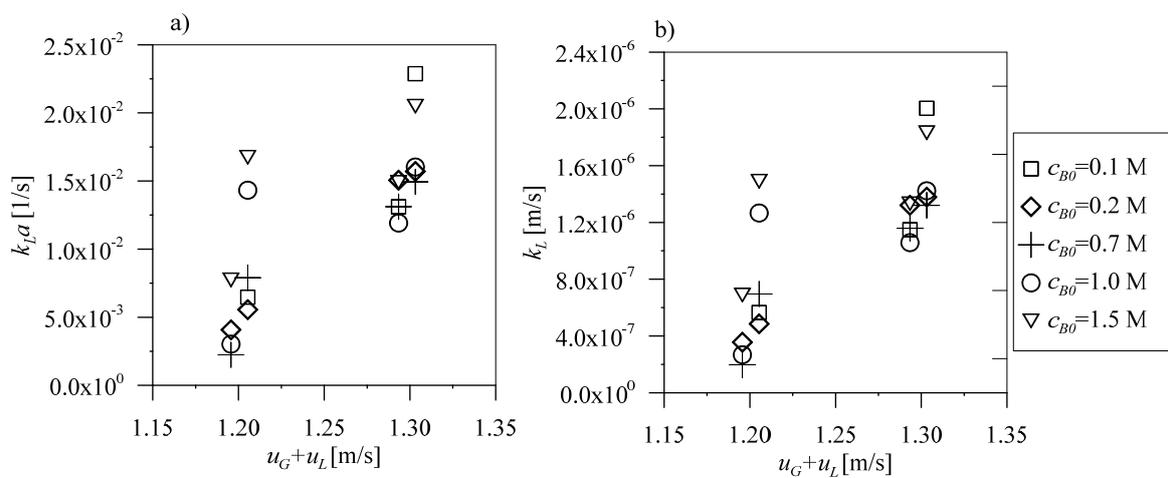


Fig. 8. k_{La} values a) and k_L values b) in function of two – phase superficial velocity

Using the estimated values of the interfacial area (Eq. (10)) values of the physical liquid-side mass transfer coefficients were determined and it was established that the condition (5) was always fulfilled (it was checked for each k_L experimental value), therefore the method is correct. Such an approach was used both for series 1 and series 2. The obtained values are presented in Fig. 8b.

A comparison of the obtained experimental values of k_L with those predicted by the Higbie’s model (Eq. (11)) and Yue et al. correlation (2007) (Eq. (13)) is presented in Fig. 9. Values of k_L were averaged for the same NaOH concentration values.

Correlation (13) predicted values of k_L close to experimental results. The differences result from the fact that Eq. (13) was suggested for the slug-annular or churn flows. Moreover, it was obtained in a straight microchannel. In the Higbie’s model, the estimation of the phase contact time is the fundamental problem. In the flow in a closed microchannel its estimation is not simple since the phase contact is connected not so much with the media superficial velocity but rather with the real phase velocities. The phase velocity may significantly diverge from u_G and u_L values due to the great importance of surface interactions (surface tension, wettability). Similar differences in the use of Higbie’s model during slug flow were obtained by Sobieszuk et al. (2011). It seems that the determination of the phase contact time is more obvious in an annular rather than in a slug flow, however, as seen in Fig. 9, the Higbie’s model should not be used for predicting mass transfer coefficients in microchannels.

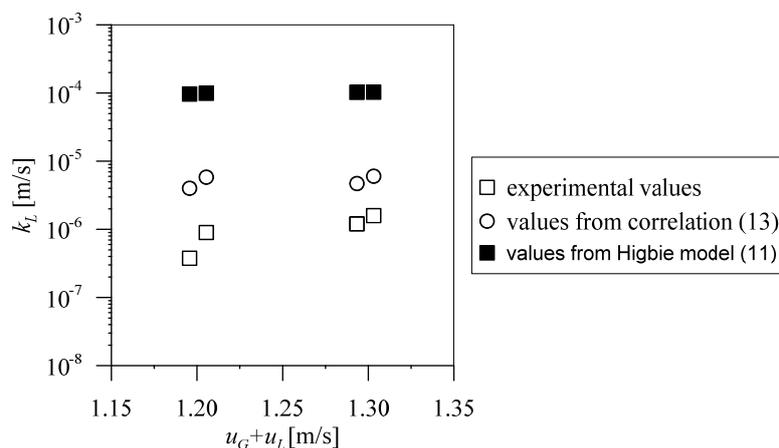


Fig. 9. The comparison of experimental k_L values with values predicted by Higbie's model (11) and correlation (13)

4. CONCLUSIONS

There are few reports in the literature concerning mass transfer in the regime of annular flow in microreactors. The paper presented the investigation on CO₂ absorption rate in NaOH aqueous solutions. The microreactor had a meandered shape with a square cross-section. The investigations were carried out in annular flow regime. Values of $k_L a$ and k_L were determined. Interfacial area was determined using the correlation (8) for the liquid film thickness in a microchannel of a rectangular cross-section. The obtained results for the liquid-side mass transfer coefficients were compared with the values resulting from the Higbie's model and a correlation developed by Yue et al. (2007), the latter obtained in a straight microchannel in slug-annular and churn flows. Similar values of k_L coefficients to those predicted from Yue et al. (2007) correlation were established. The problem with correct determination of contact time between the phases proves that the Higbie's model, similarly to a slug flow, should not be used in an annular regime in microchannels either.

SYMBOLS

a	interfacial area per unit volume, m ² /m ³
b	stoichiometric coefficient
Ca	capillary number, - ($Ca = u_{TP} \mu_L / \sigma$)
c_A	concentration of CO ₂ in gas phase, mol/m ³
c_{B0}	initial concentration of NaOH in liquid phase, mol/m ³
d_H	hydraulic diameter of microchannel, m
D_A	diffusivity of CO ₂ in liquid phase, m ² /s
D_B	diffusivity of OH ⁻ in liquid phase, m ² /s
E	enhancement factor
H	Henry's constant, mol / (m ³ · Pa)
k_L	liquid - side mass transfer coefficient, m/s
$k_L a$	volumetric liquid - side mass transfer coefficient, 1/s
k_2	rate constant of reaction CO ₂ with ions OH ⁻ , m ³ /(mol · s)
L	microchannel length, m
N_A	absorption rate, mol/s
p_A	CO ₂ partial pressure, Pa
Q	flow rate, m ³ /s

Re_G	Reynold number for gas phase, - ($Re_G = u_G \rho_G d_H / \mu_G$)
Re_L	Reynold number for liquid phase, - ($Re_L = u_L \rho_L d_H / \mu_L$)
R	gas constant, J/(mol·K)
r	radius of liquid flow cross-section, m
Sc	Schmidt number, - ($Sc = \mu_L / (\rho_L \cdot D_A)$)
Sh	Sherwood number, - ($Sh = k_L \cdot d_H / D_A$)
T	temperature, K
u	superficial velocity, m/s
V	microchannel volume, m ³

Greek symbols

δ	liquid film thickness, m
ρ	density, kg/m ³
μ	dynamic viscosity, Pa·s
σ	surface tension, N/m
τ	contact time, s

Superscripts

*	with chemical reactions
AV	average

Subscripts

I	inlet
2	outlet
G	gas
L	liquid
TP	two - phase

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